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Allocation Candidates are
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3

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Mass of sodium hydroxide = 3.5 g
Mass of biodiesel produced =

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811.0 g The relative molecular mass of the oil used by the student is 885.6. Calculate the amount (in moles) of the oil and the methanol used, and hence the amount (in moles) of excess methanol..... [3] (This question continues on the following page)

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$(l)C + 3 CH_3OH (l) (l) + 3 H_2C=O \dots$

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- 3 - Turn over

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(Question A1 continued) (d)

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*Pentan-2-one has the following
mass spectrum. Relative
abundance 100 80 60 40 20*

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*coefficient of 3 on stearic acid; [3]
Accept displayed or condensed
formulas for molecules (b) (i) both
have first double bond on C9 with
carbon / linoleic has an $\omega - 6$ C=C*

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1 Question 2 Question 3 Question
4 Question 5 Did you know?
Multiple choice questions from
Paper 1 are worth 1 mark, no
matter their ...*

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3 2 PO B. NH 3 HCO 3 PO 4 3 C.
NH 4 HCO 3 2 PO 4 2 D. NH 4
HCO 3 PO 4 3 12. What happens
when magnesium metal reacts

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with chlorine gas? A. Each magnesium atom loses two electrons and each chlorine atom gains two electrons. B. Each magnesium atom gains one electron and each chlorine atom loses one electron. C.

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Scheme [Cd²⁺(aq)] / mg dm⁻³
correct labelling showing
concentration (of cadmium ions)*

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*with units and appropriate scales;
Do not penalize for use of smaller
but consistent divisions. correct
plotting and straight line through
the origin; concentration of Cd²⁺
(from graph) 0.36(mg dm⁻³) 3; [3]
Allow answer in range 0 ...*

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any of the above papers? I'm
mostly looking for paper 2 but
any from May 2014 will do. 0*

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3: Reference code:
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Determine and State ...*

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