

Reaction Rates And Equilibrium Practice Problems Answers

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A very small equilibrium constant (<0.001) means very little product will be present at equilibrium (or that the reverse reaction goes to completion). Changing the initial concentrations of reactants or products does NOT change the equilibrium constant. The equilibrium concentrations of reactants and products will change but Keq will not change.

[Equilibrium Practice Problems: using equilibrium constants ...](#)

In chemistry, and in physics, a dynamic equilibrium exists once a reversible reaction occurs. Substances transition between the reactants and products at equal rates, meaning there is no net change. Reactants and products are formed at such a rate that the concentration of neither changes. It is a particular example of a system in a steady state.

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